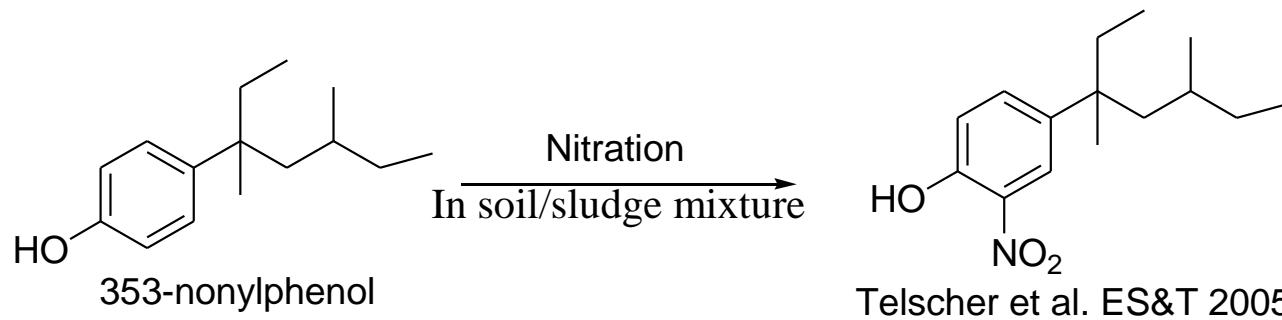
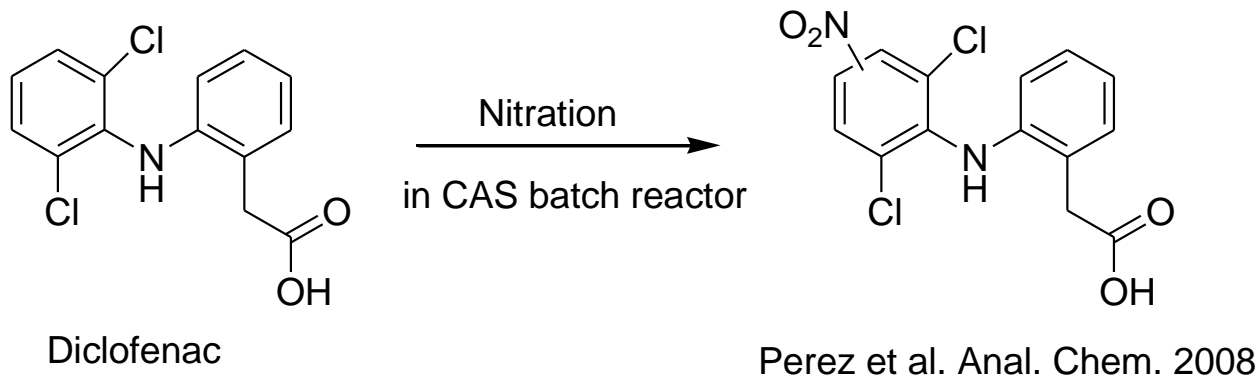
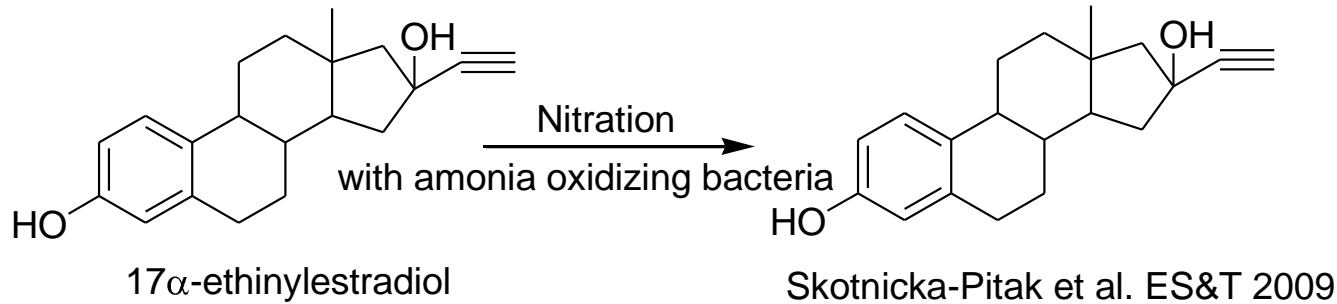


Nitration processes of acetaminophen in nitrifying activated sludges

Serge Chiron, Marseille University

E.Gomez and H.Fenet, Montpellier University

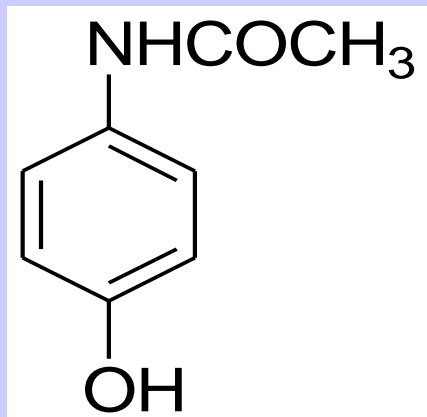
An unexpected biotransformation pathway: nitration



Objectives

- **To investigate the nitration mechanisms of phenolic compounds in nitrifying activated sludge.**
- **At different scales: field-, batch- and molecular-scale experiments.**
- **Acetaminophen (paracetamol) as a probe compound.**

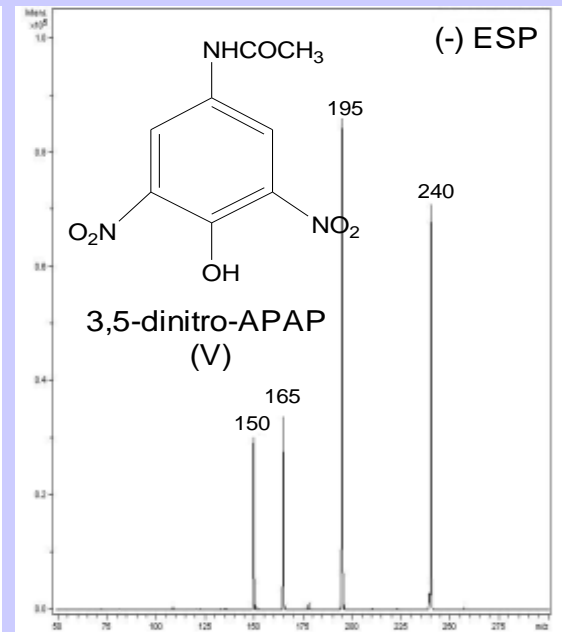
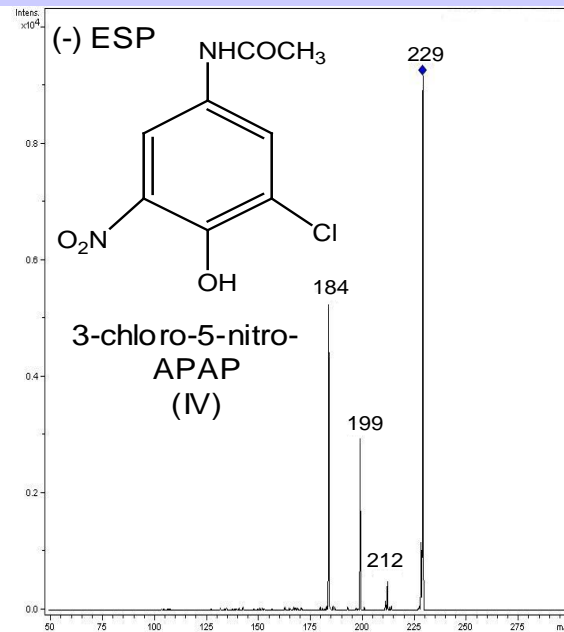
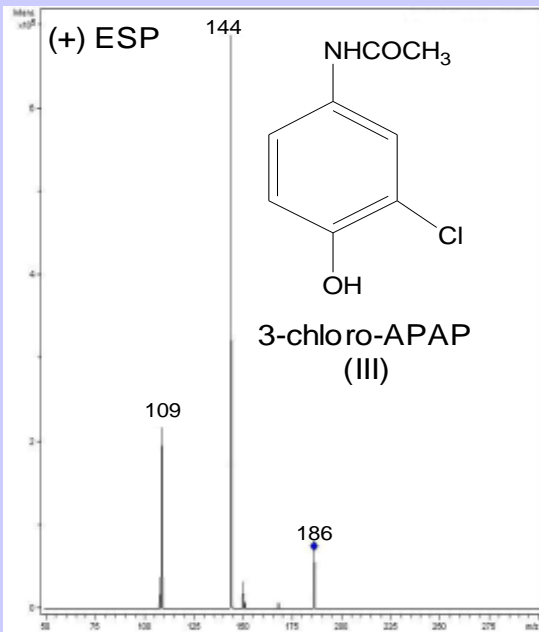
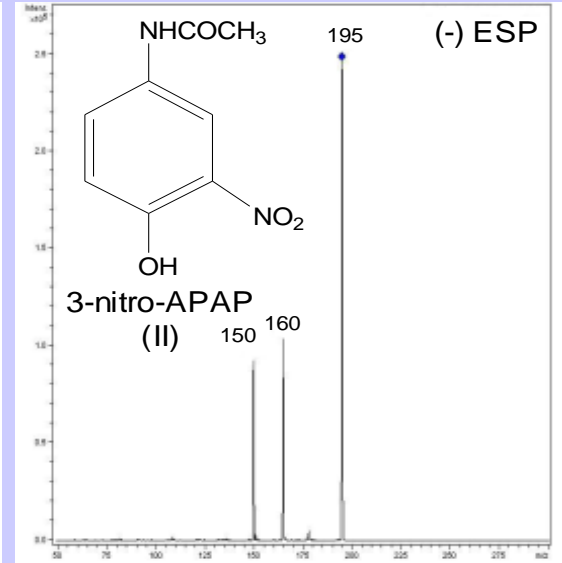
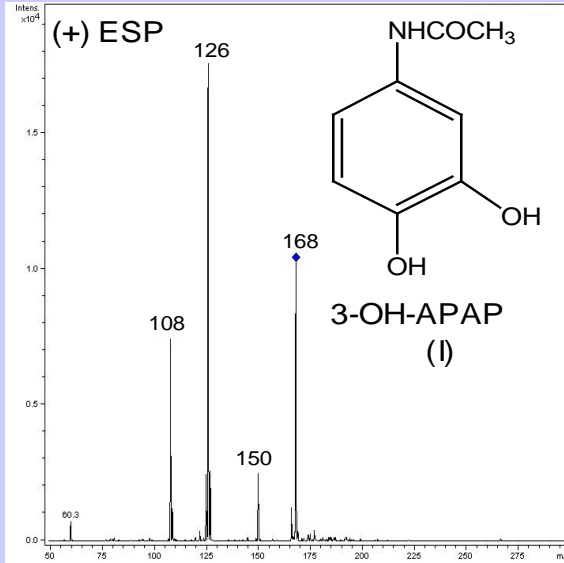
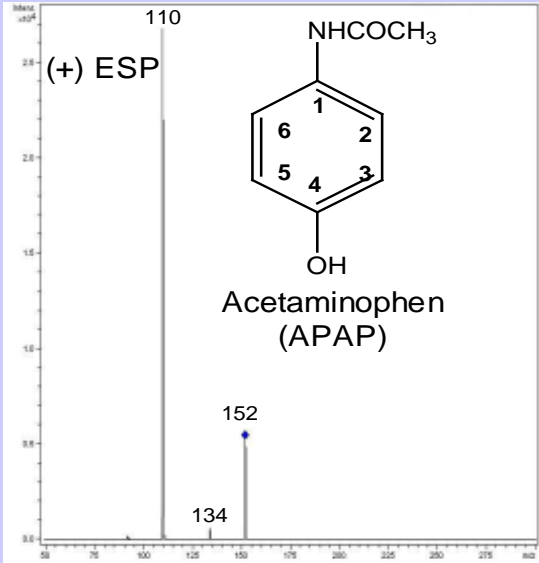
Why paracetamol ?



$\text{pK}_a = 9.4$
 $\log K_{ow} = 0.45$
 $S_w > 1\text{g/L}$

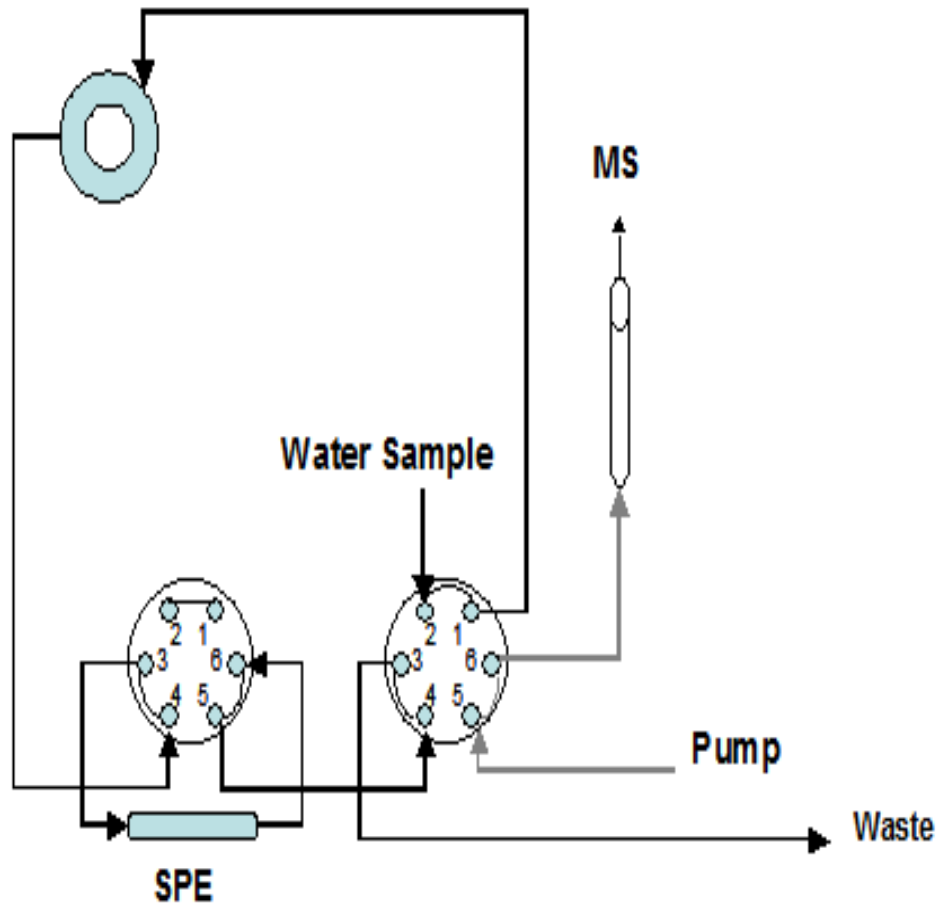
- **Readily prone to nitration**
- **High occurrence in WWTPs (1-10 $\mu\text{g/L}$ range)**
- **Nitrated derivatives can be easily synthesized**

Field studies: Targeted compounds



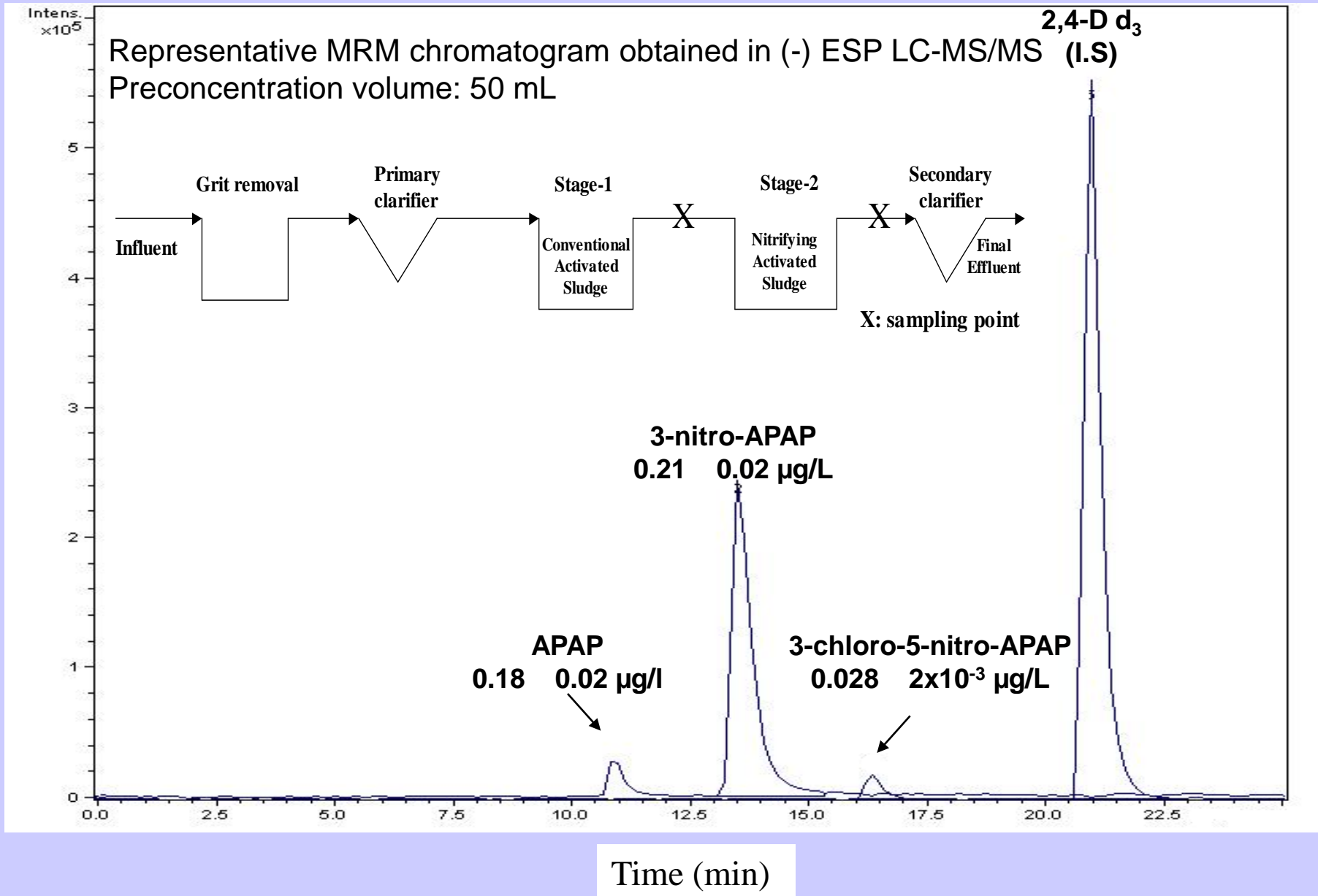
Analytical methodology

On-line SPE-LC-MS/MS



Compound	Recoveries (%)
3-OH-APAP	55 (\pm 12)
APAP	75 (\pm 8)
3-nitro-APAP	78 (\pm 8)
3-chloro-APAP	86 (\pm 6)
3-chloro-5-nitro-APAP	93 (\pm 4)
3,5-dinitro-APAP	95 (\pm 5)

Aix-en-Provence WWTP effluent analysis



Field data (24 h composite samples)

	Sampling month	APAP	3-OH-APAP	3-chloro-APAP	3-nitro-APAP	3-chloro-5-nitro-APAP
Stage-2 influent	Oct. 2008	3.45 ± 0.21	0.96 ± 0.11	0.24 ± 0.02	n.d	n.d
	Nov. 2008	5.35 ± 0.32	1.44 ± 0.17	0.85 ± 0.04	n.d	n.d
	Dec. 2008	6.75 ± 0.41	1.86 ± 0.22	0.76 ± 0.04	n.d	n.d
Stage-2 effluent	Oct. 2008	0.19 ± 0.02	n.d	n.d	0.18 ± 0.02	0.03±1x10⁻³
	Nov. 2008	0.35 ± 0.03	n.d	n.d	0.26 ± 0.03	0.11±5x10⁻³
	Dec. 2008	0.64 ± 0.05	n.d	n.d	0.32 ± 0.03	0.09±3x10⁻³

Batch experiments

Experimental conditions:

$[\text{MLSS}] = 2.5 \text{ g/L}$

$\text{pH} = 7\text{-}7.5$

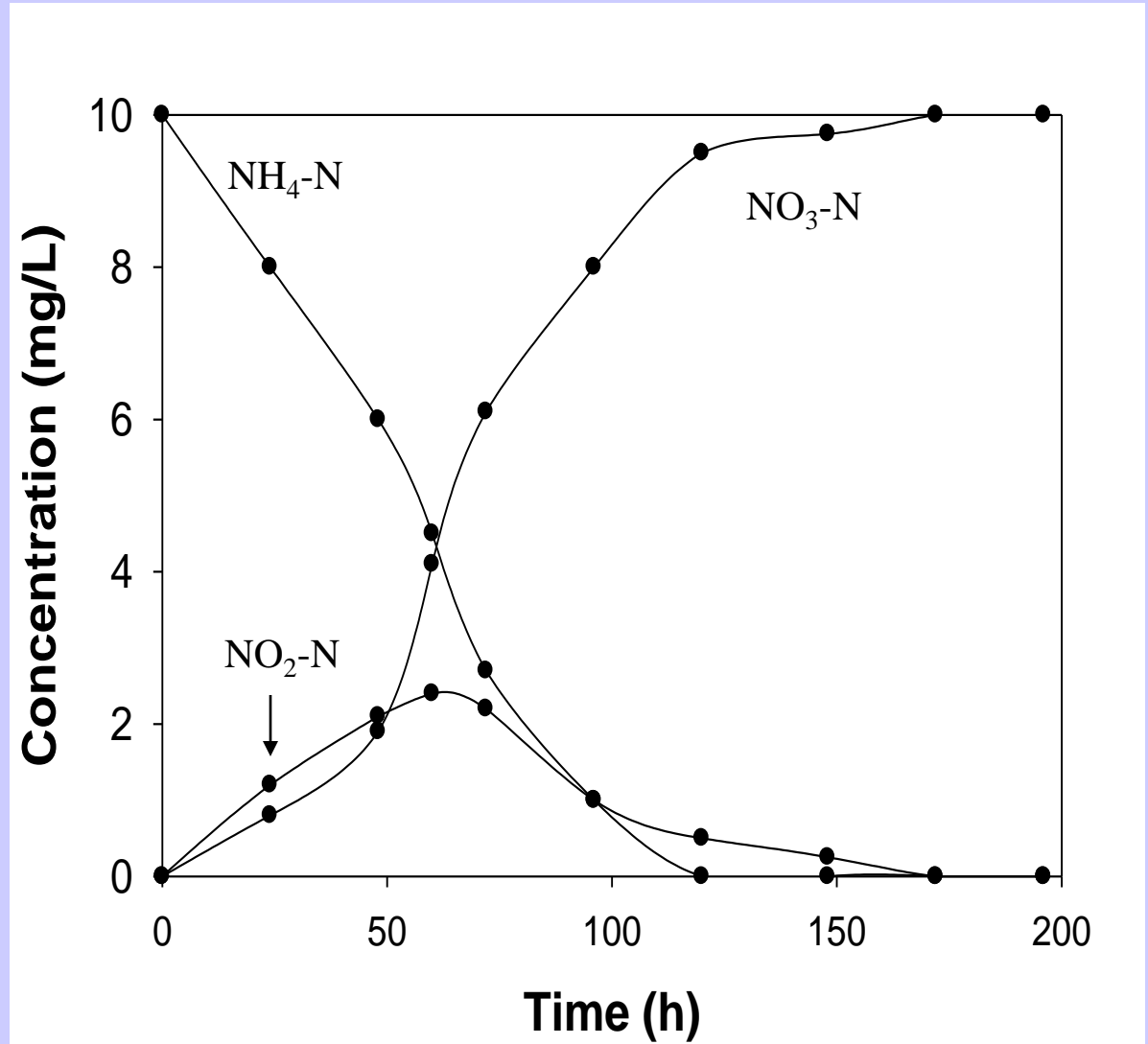
$T = 25 \text{ C}$

$[\text{O}_2] > 3 \text{ mg/L}$

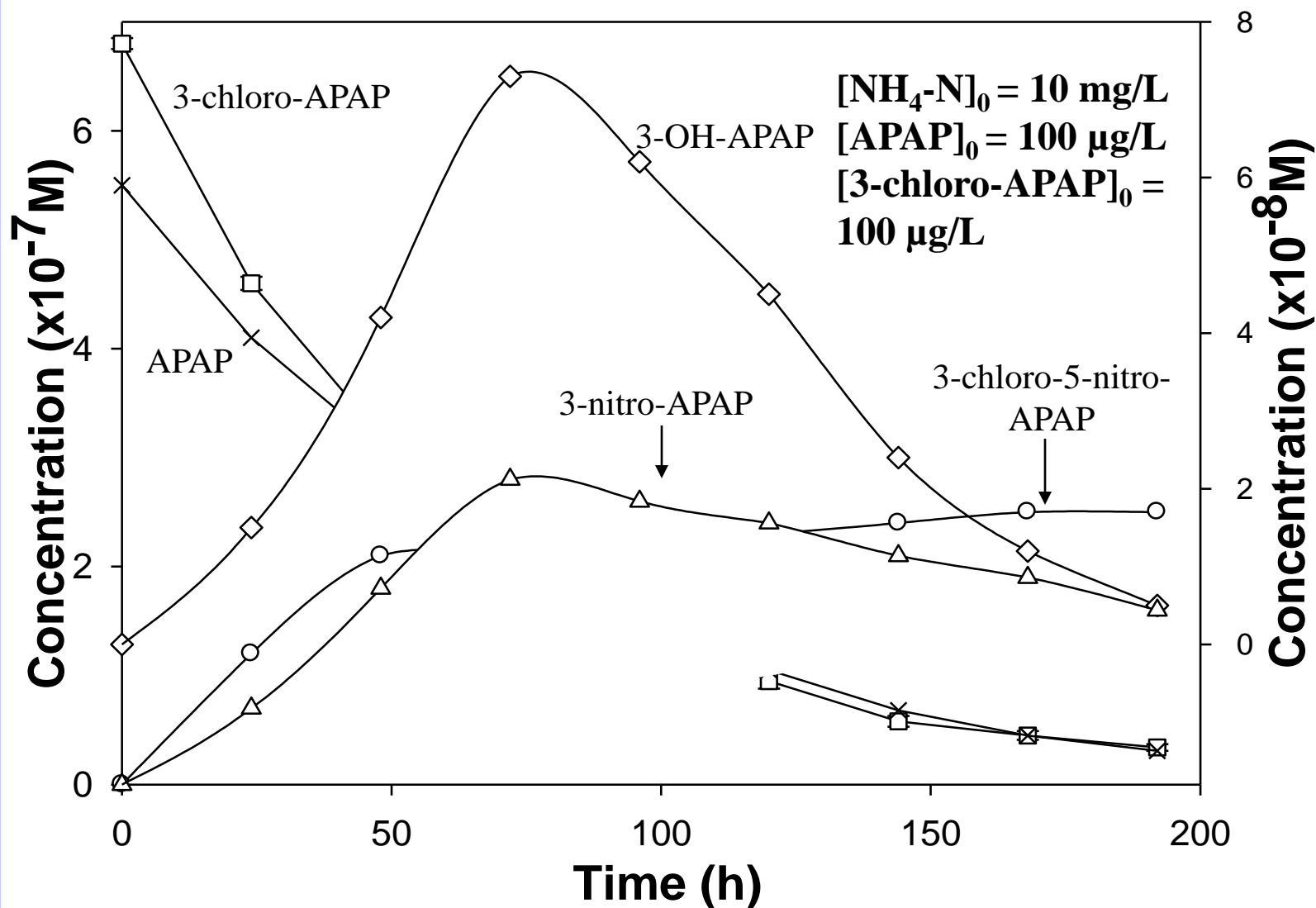
$[\text{NH}_4^+] = 10 \text{ mg/L}$

$[\text{APAP}]_0 = 100 \text{ }\mu\text{g/L}$

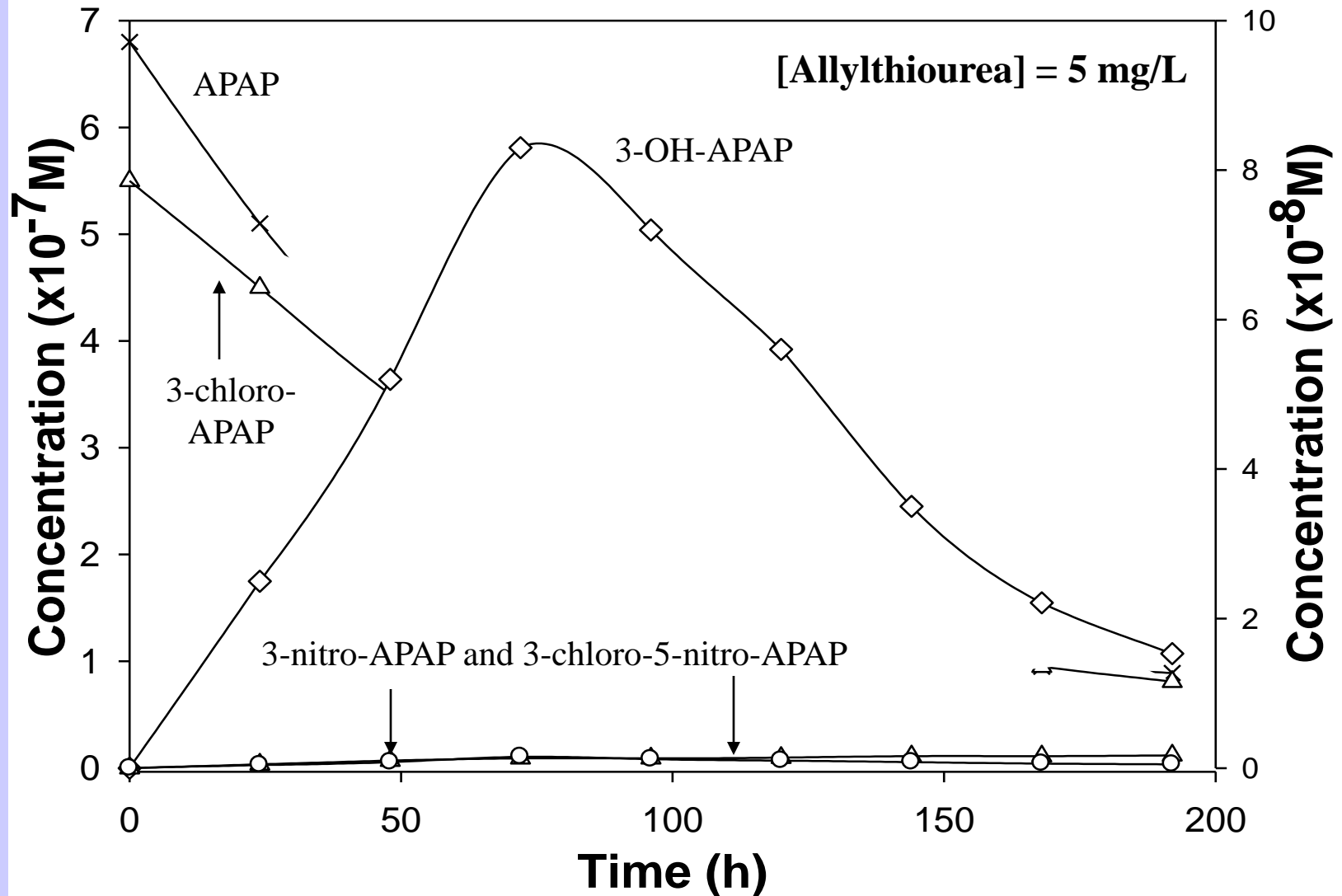
$[\text{3-chloro-APAP}]_0 = 100 \text{ }\mu\text{g/L}$



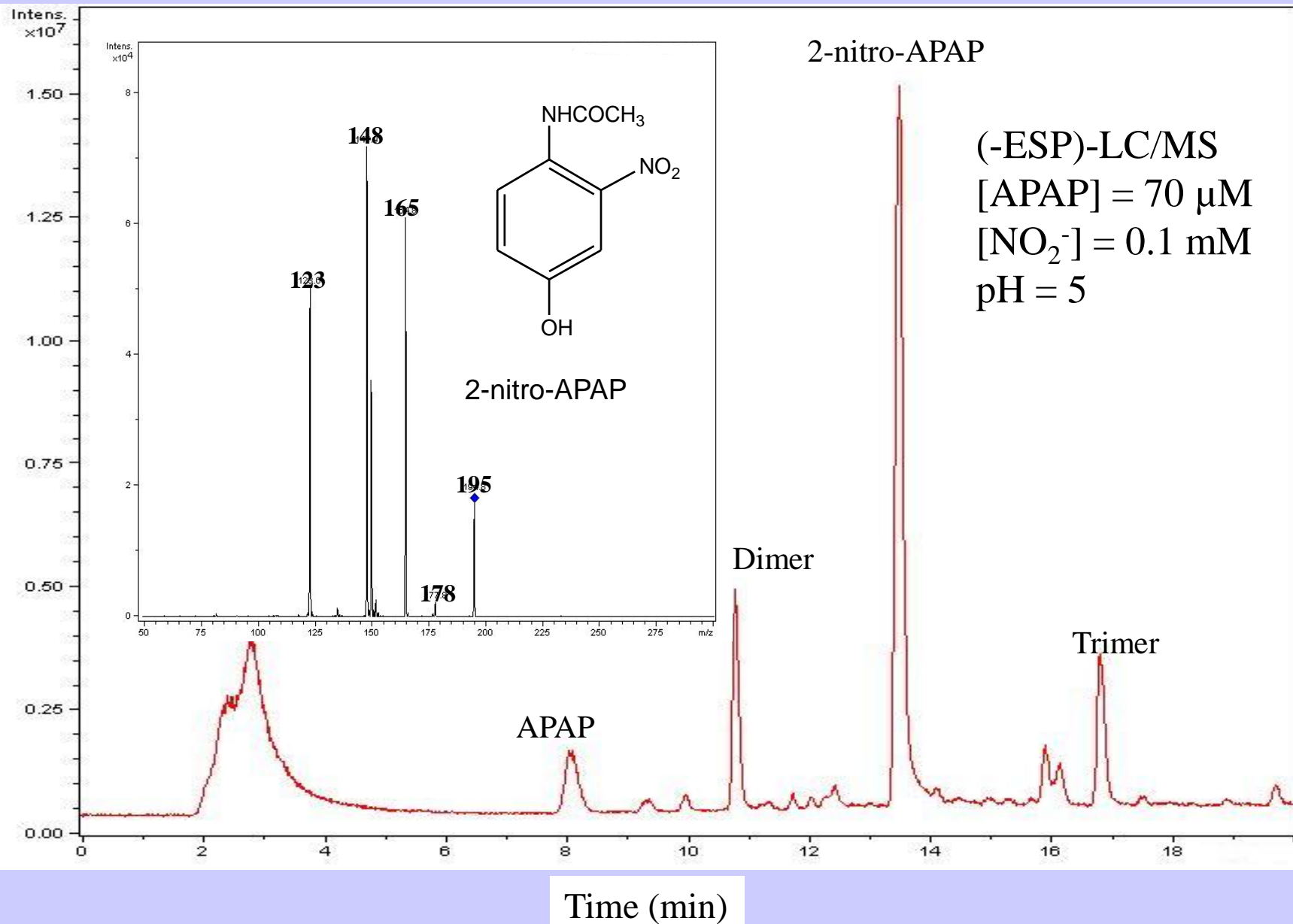
Batch 1 experiment: no nitrification inhibition



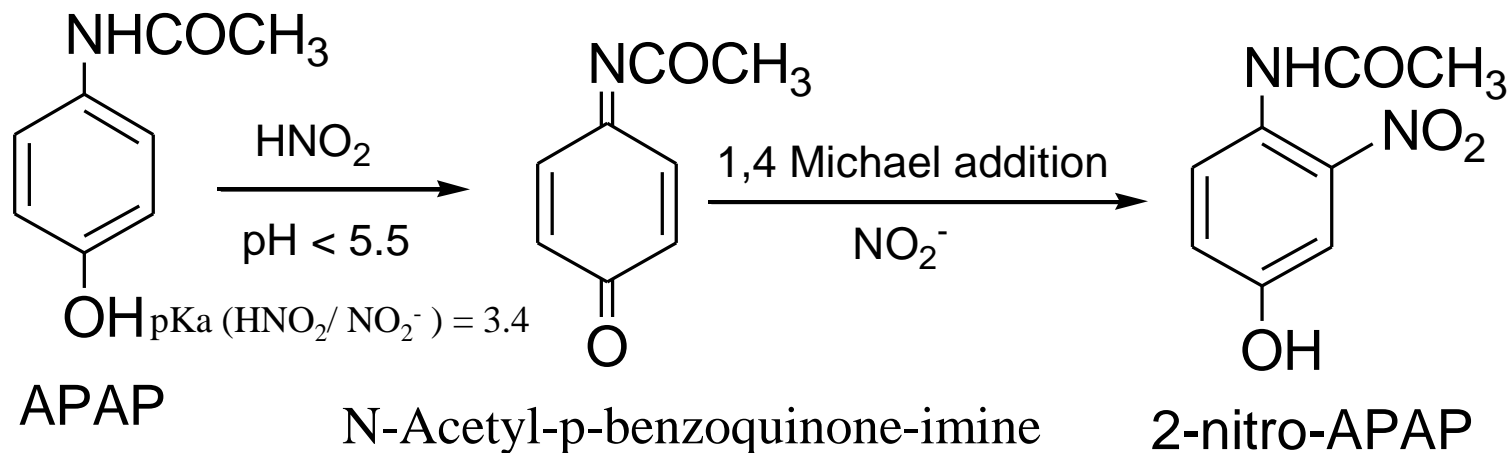
Batch 2 experiment : ammonia oxygenase inhibition



Reaction of APAP with HNO₂



Proposed mechanism of 2-nitro-APAP formation



Matsumo et al. Chem. Pharm. Bull. 1989

Reactivity of APAP with horseradish peroxidase

(-ESP)-LC/MS

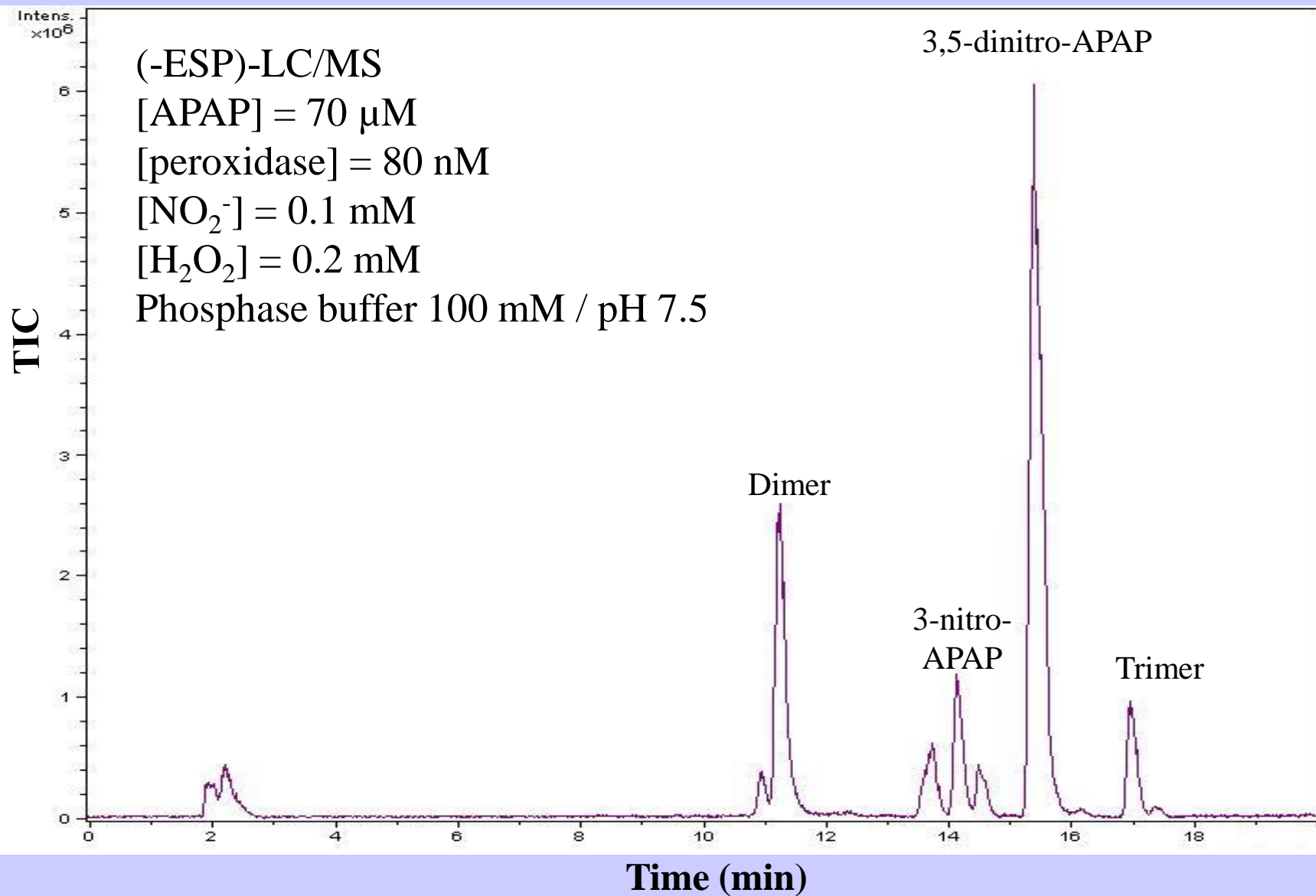
[APAP] = 70 μM

[peroxidase] = 80 nM

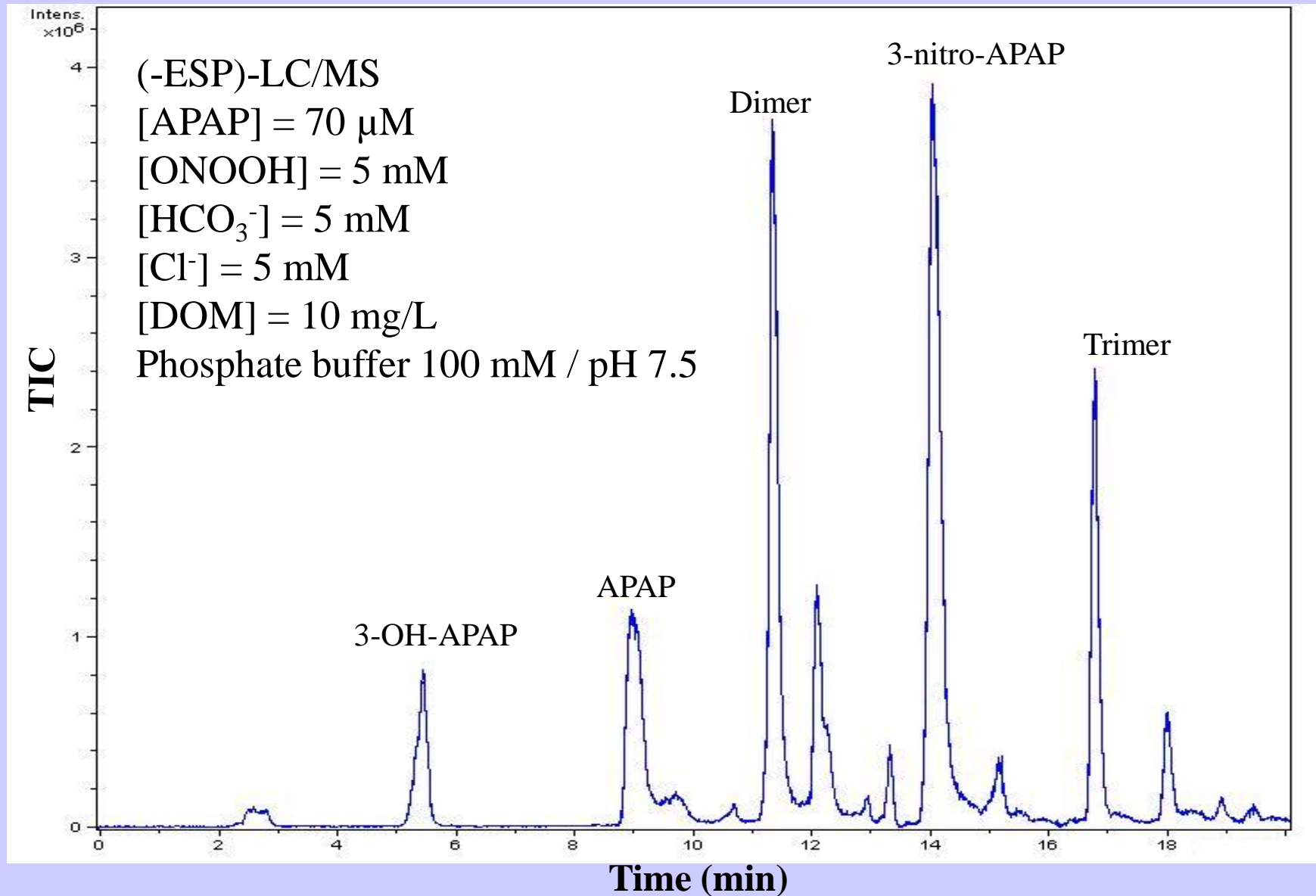
[NO₂⁻] = 0.1 mM

[H₂O₂] = 0.2 mM

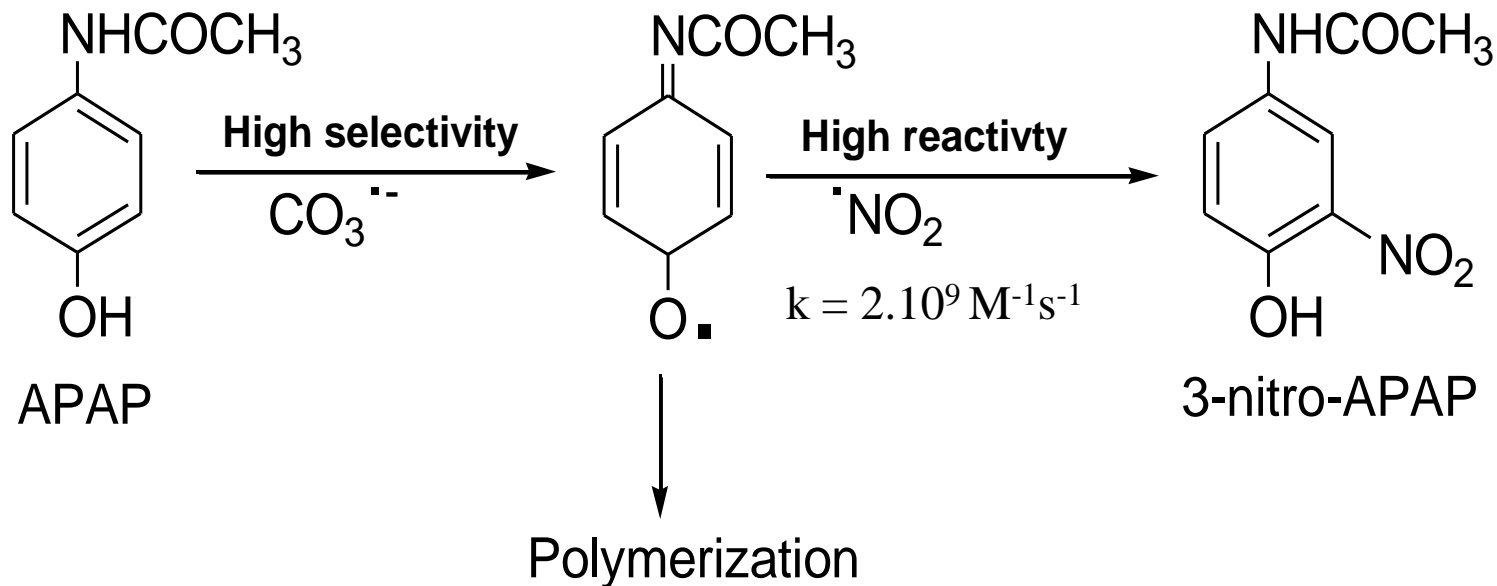
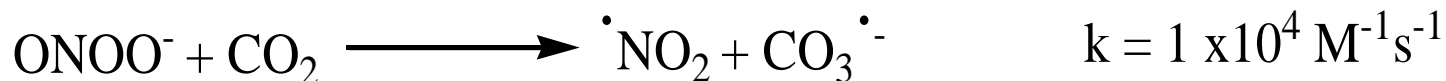
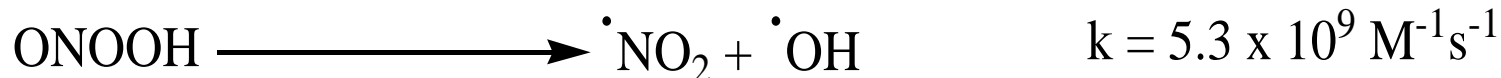
Phosphate buffer 100 mM / pH 7.5



Reactivity of APAP with peroxyntirite



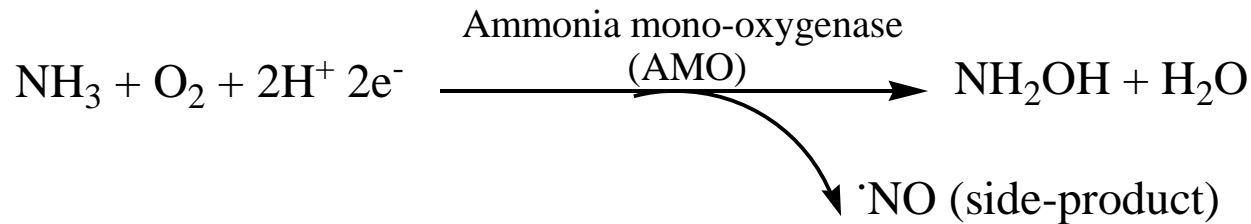
Proposed mechanism of 3-nitro-APAP formation



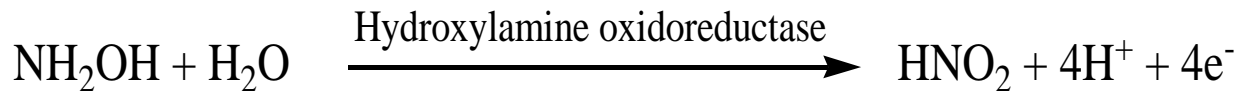
Where is peroxyxynitrite coming from?

NH_4^+ oxidation by ammonia oxidizing bacteria (AOB)

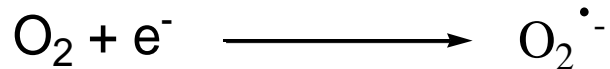
- First step



- Second step



- Third step

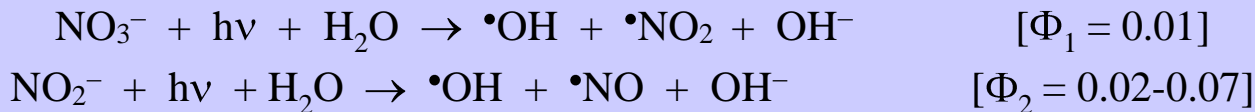


Conclusions

- **Formation of nitro-APAP derivatives with an environmental profile more worrisome than APAP.**
- **Nitration is probably linked to $\cdot\text{NO}$ generation by nitrifying bacteria.**
- **This result must be validated with other phenolic pollutants (i.e. bisphenol A, nonylphenol).**

Nitration processes in the environment

- Photonitration

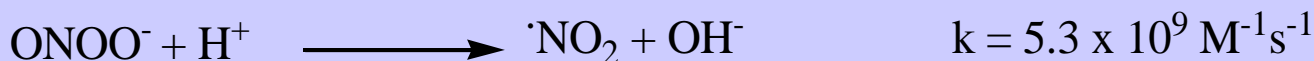


- Thermal nitration by HNO_2 (in the dark)

pKa ($\text{HNO}_2 / \text{NO}_2^-$) = 3.4 relevant at pH < 5.5

- Bionitration by peroxidases in presence of NO_2^- and H_2O_2

- Bionitration by ONOOH (peroxynitrite)



Expected biotransformation pathways

